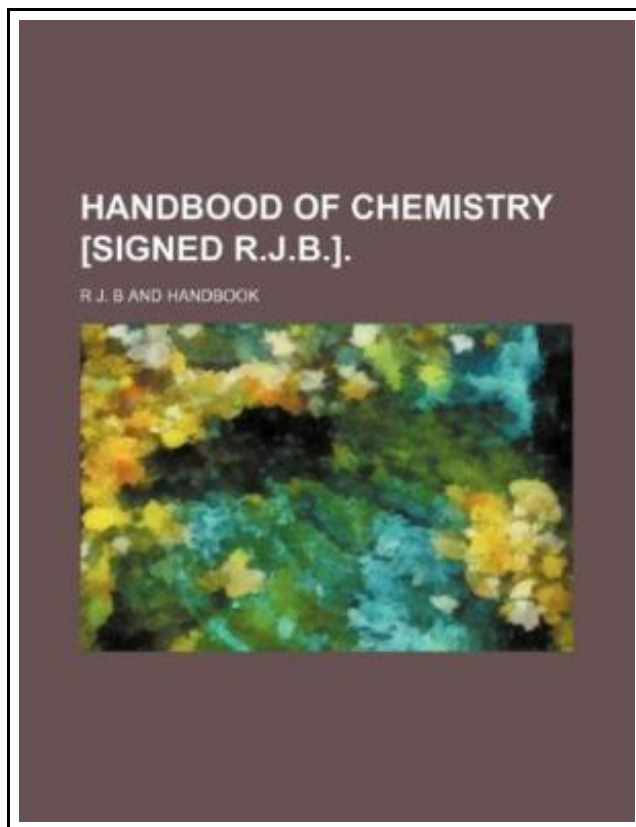


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1868 Excerpt: . gas over the pneumatic trough. The decomposition is-- CaO , $\text{CO}_2 + \text{HCl} = \text{CaCl} + \text{HO} + \text{CO}$, . The usual test for this gas is lime-water. When limewater is mixed with carbonic acid, carbonate of lime is formed, which is known by its milky colour. The presence of carbonic acid in the air maybe proved by placing a dish containing lime water in a room, and leaving it for a few hours. It will gradually turn milky, on account of its slow conversion into carbonate of lime. Carbonic acid is also found in water, and is of great use to plants, since it provides them with the carbon necessary for their growth. The carbonates of lime and magnesia are found in the water of wells and rivers, but they are dissolved by the carbonic acid, and thus rendered invisible. Since it is found that carbonic acid is present in the air, the question is--how does it come there? There are many ways, but the chief are by means of combustion, respiration, fermentation, putrefaction, and decay. Bach of these must be taken up so as to show in what manner it contributes carbonic acid to the air. Combustion.--Place a lighted taper in a glass jar (see fig. 20), and then test the gas left in the jar. It will be found to contain carbonic acid, which is made evident by the lime water, used as a test, being changed into the milky carbonate of lime. Thus it is manifest that wherever combustion is going...



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